

Relationship between K_{eq} and ΔG°

$$\Delta G^{\circ} = - [1372] \lg_{10} K_{eq}$$

<u>Products</u> <u>Reactants</u>	K_{eq}		\lg_{10}	ΔG° cal/mole* [$\lg_{10} \times -1372$]	
1/1000	.001	10^{-3}	-3	+ 4116	[R] > [P]
1/100	.01	10^{-2}	-2	+ 2744	
1/10	.1	10^{-1}	-1	+ 1372	
1/1	1.0	0	0	0	
10/1	10	10^{+1}	+1	- 1372	[P] > [R]
100/1	100	10^{+2}	+2	- 2744	
1000/1	1000	10^{+3}	+3	- 4116	