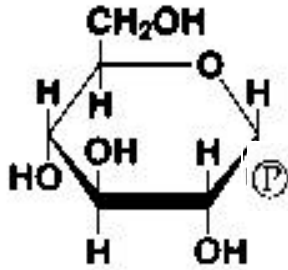


glucose 1 P

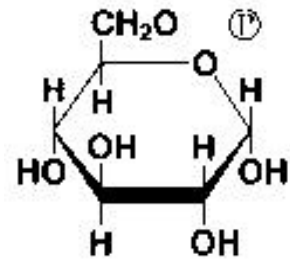


phospho gluco mutase



$$\Delta G^0 = -1.74 \text{ Kc/m}$$

glucose 6 p



initiate reaction @ 0.2 molar concentrations

@ equilibrium measure Keq

$$\frac{[P]}{[R]} = \frac{G6P}{G1P} = \frac{0.19}{0.01}$$

$$Keq = 19$$

$$\Delta G^0 = -1,372 \log_{10} Keq$$

$$= -1,372 \log_{10} 19$$

$$= -(1,372) (1.2788)$$

$$= -1,754 \text{ cal/mole}$$

Exergonic & Spontaneous

@ 250C ... -RT ln Keq $\Delta G^0 = -(2.0) (298) (2.303) \lg_{10} Keq = -[1374] \lg_{10} Keq$