#	SYMBOL	ELEMENT	ELECTRONEGATIVITY
> greater the Value the more Electronegative			
1	O	Oxygen	3.44
2	Cl	Chlorine	3.16
3	N	Nitrogen	3.04
4	Br	Bromine	2.96
5	S	Sulfur	2.58
6	C	Carbon	2.55
7	Se	Selenium	2.55
8	H	Hydrogen	2.20
9	P	Phosphorus	2.19
10	As	Arsenic	2.18
11	Mo	Molybdenum	2.16
12	B	Boron	2.04
13	Ni	Nickel	1.91
14	Si	Silicon	1.90
15	Cu	Copper	1.90
16	Co	Cobalt	1.88
17	Fe	Iron	1.83
18	Zn	Zinc	1.65
19	Al	Aluminum	1.61
20	Mn	Manganese	1.55
22	Mg	Magnesium	1.31
23	Ca	Calcium	1.00
25	Li	Lithium	0.98
26	Sr	Strontium	0.95
27	Na	Sodium	0.93
28	K	Potassium	0.82

Electronegativity,

is the tendency for an atom of a given chemical element to attract shared electrons when forming a chemical bond.

An atom's electronegativity is affected by both its atomic number and the distance at which its valence electrons reside from the charged nucleus.